

Design and optimisation of a pulsed CO₂ laser for laser ultrasonics

A. Forbes¹, L.R. Botha¹, N. du Preez² and T. Drake³

¹CSIR National Laser Centre, ²SDI (Pty) Ltd., ³Lockheed Martin Aerospace Company

Presented at

CSIR R&I Conference

Dr Andrew Forbes

Senior Scientist

February 2006

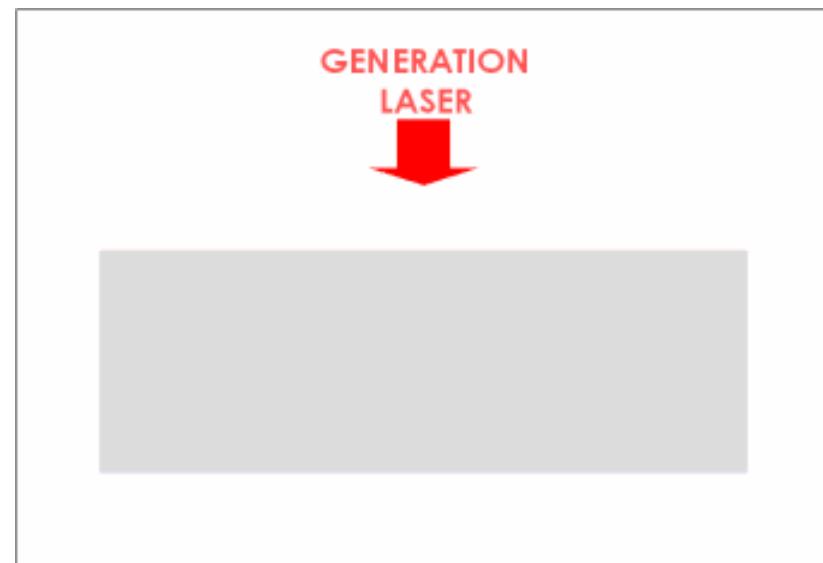


Contents

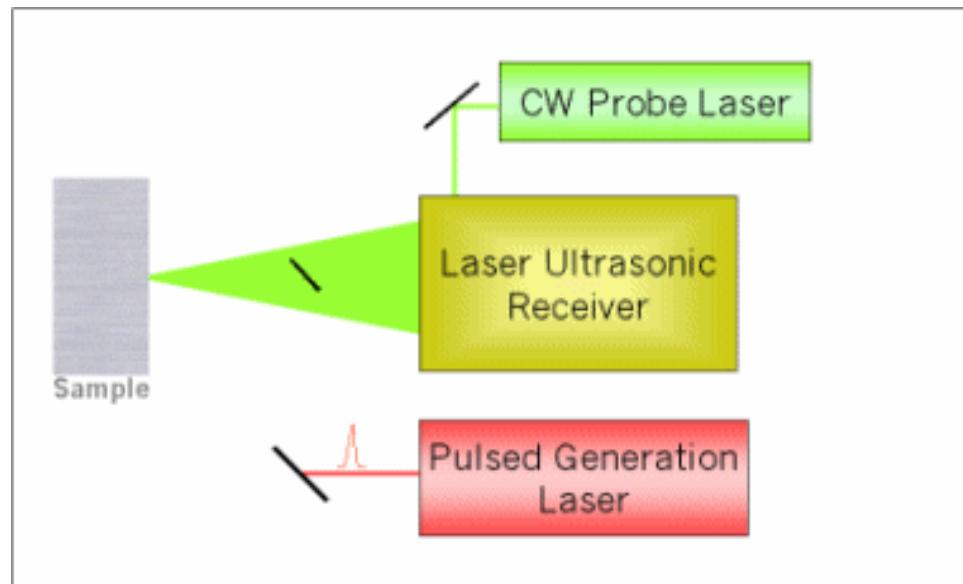
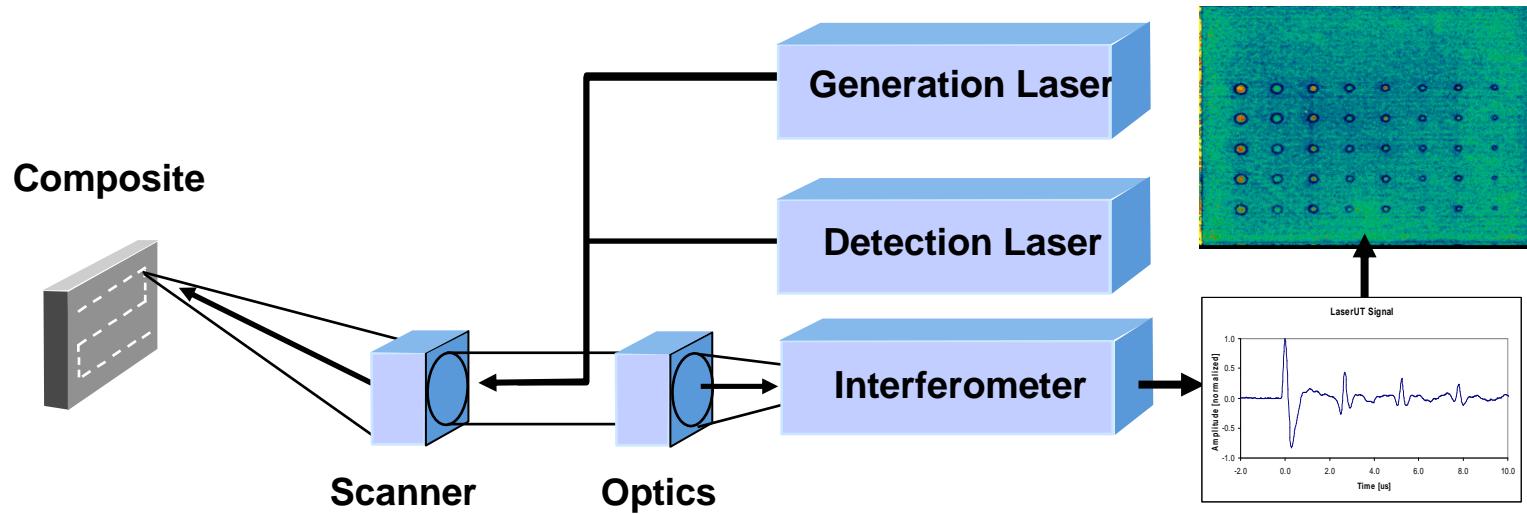
- **Laser ultrasonics**
What is it, and how does it work?
- **Optimising the parameters**
Choices and consequences
- **Laser chemistry**
A physicists approach to chemistry
- **Discharge disturbances**
Bloody hell, not more problems!
- **Final system performance**
Working at last!

Laser ultrasonics – basic principles

- Optical absorption leads to thermal expansion
- Thermal expansion causes ultrasonic waves in sample
- Long pulse laser, coupled to an interferometer detects mechanical displacements produced by ultrasonic waves
- Ultrasound can be detected optically using an optical heterodyne technique



Laser ultrasonics – basic principles



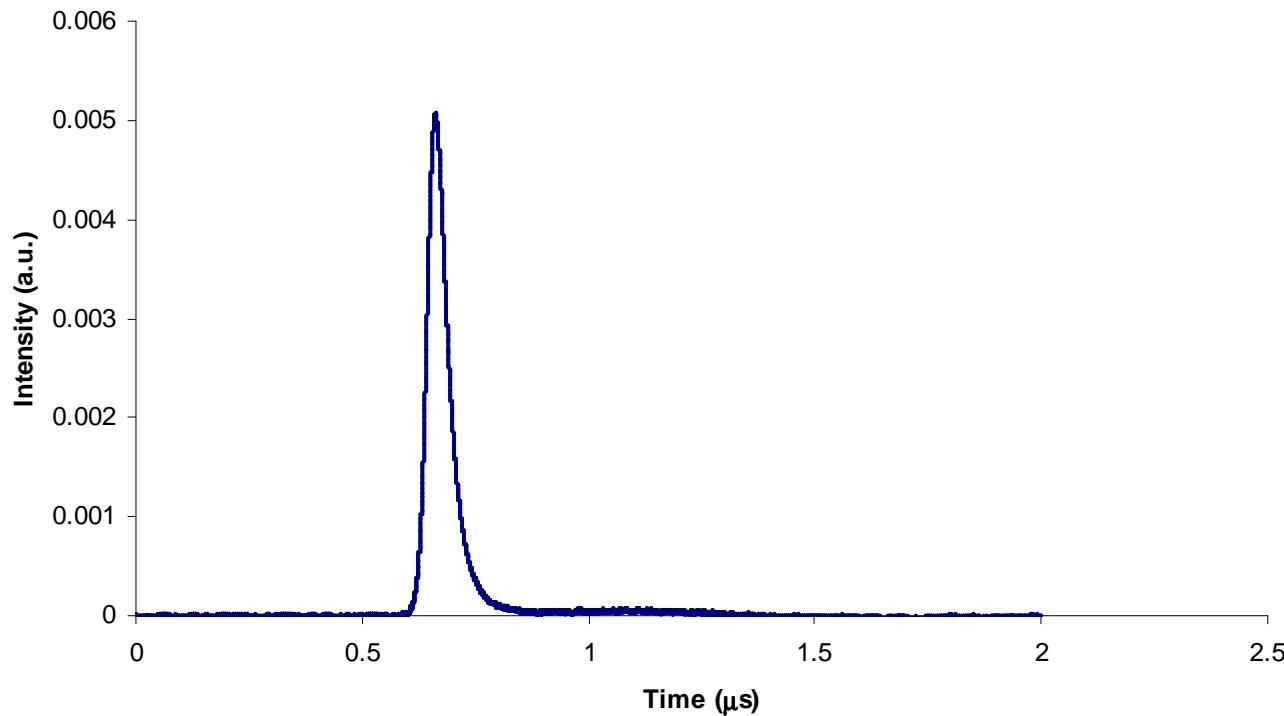
Optimising the parameters

How do you design a laser for this application?



Objective

- Optimise the parameter set
 - Maximise energy,
 - Maximise repetition rate,
 - Maximise acoustic signal efficiency. → Short time pulses



Laser parameters

- Choices and consequences

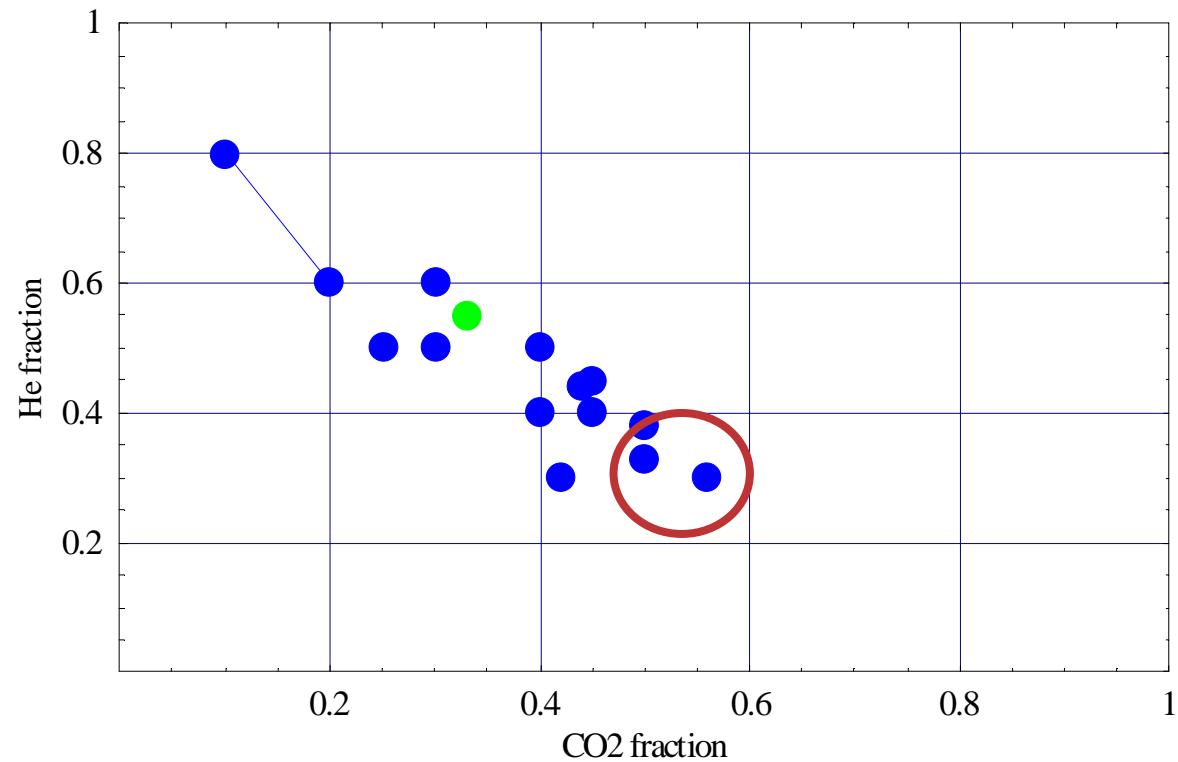
Pulse Envelope	<i>Factors affecting this are:</i>		
OC Reflectivity	HV	N ₂	CO ₂

Energy	<i>Factors affecting this are:</i>		
OC Reflectivity	HV	N ₂	CO ₂

Stability	<i>Factors affecting this are:</i>			
Electrodes	HV	Acoustics	Pre-ionisation	Catalysts

Lifetime	<i>Factors affecting this are:</i>		
OC Reflectivity	HV	Catalysts	Gas Mix

Gas influences

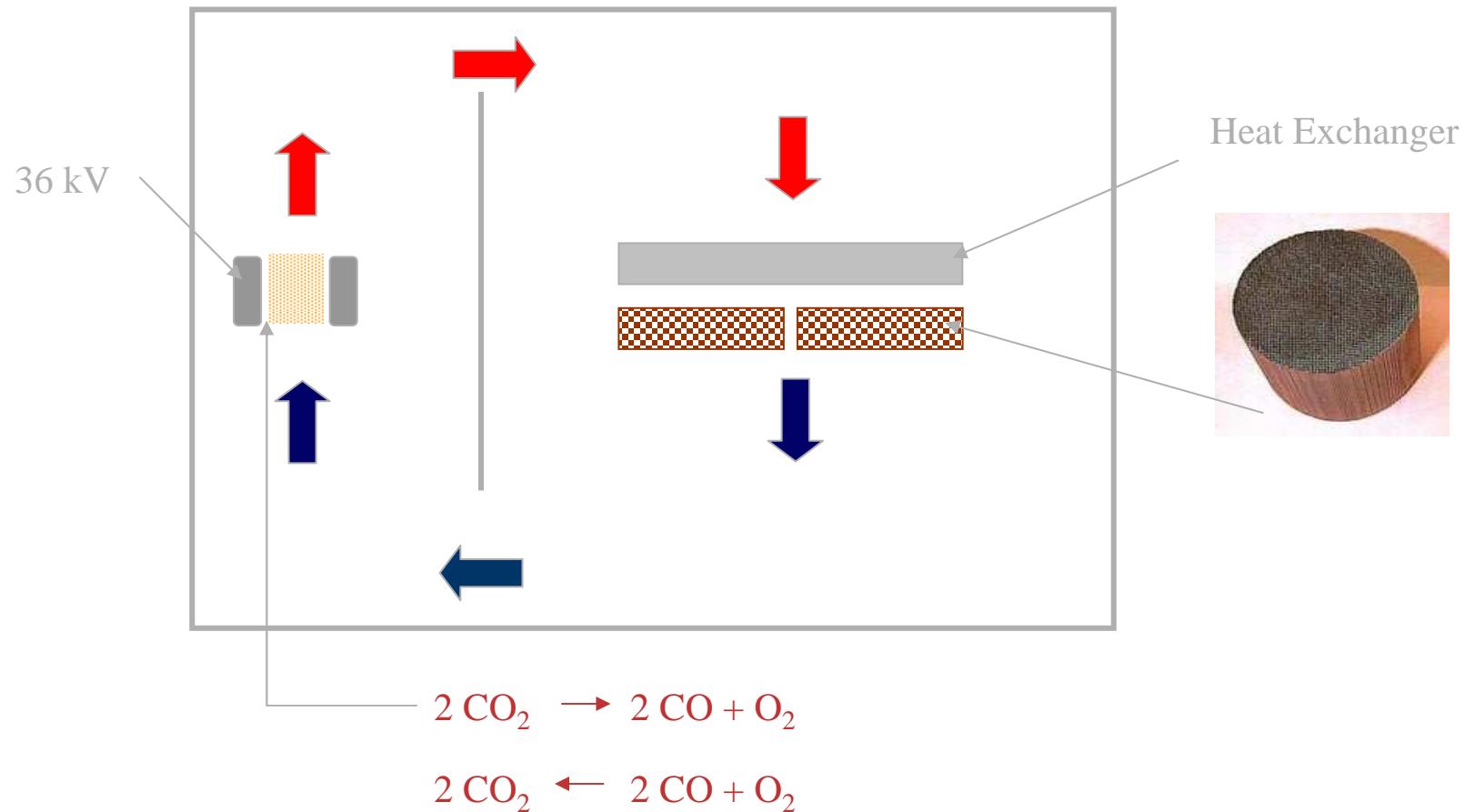


Laser chemistry

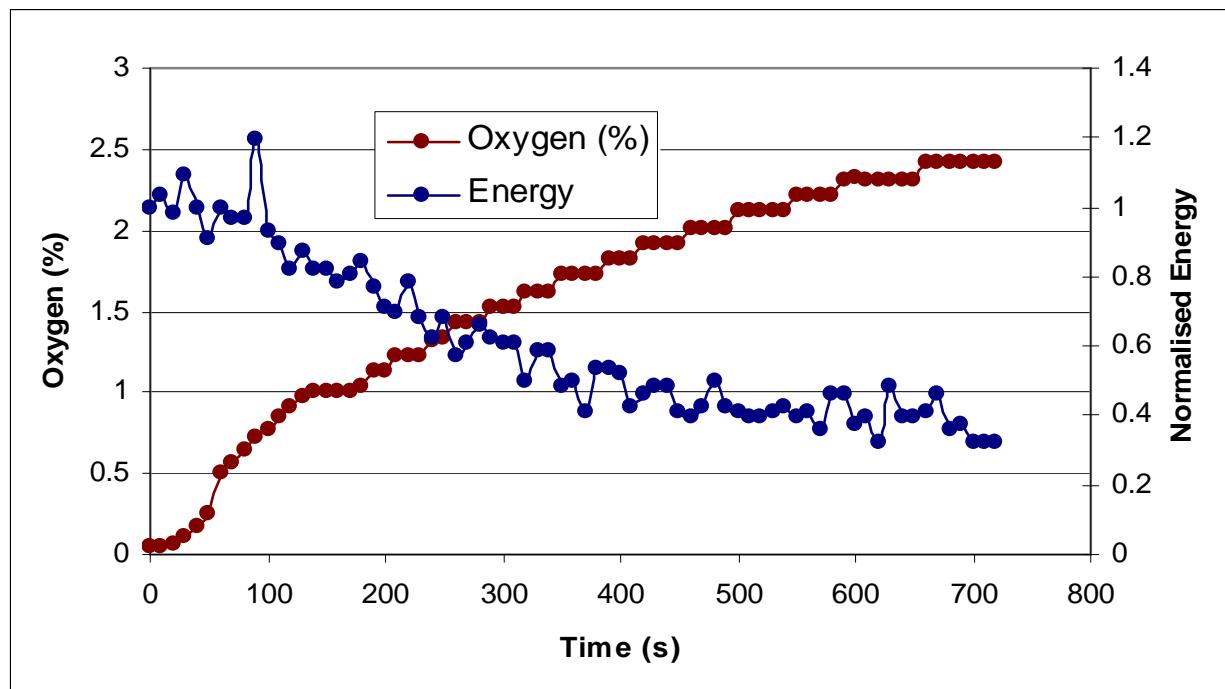
Impact of gas mix on laser chemistry



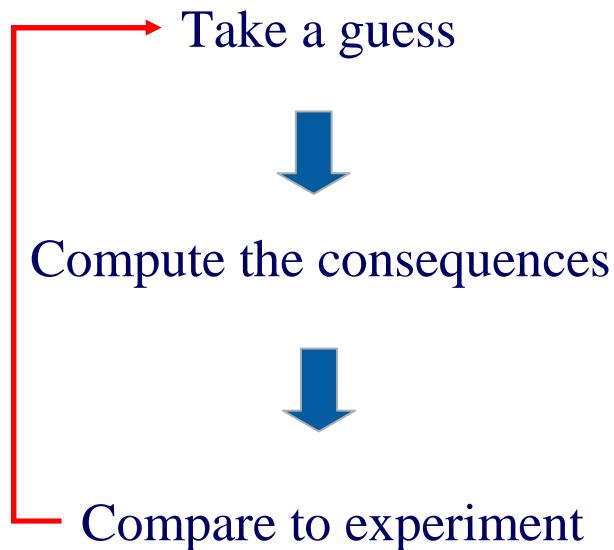
Problem statement



Typical problem



Approach



But, cannot determine rate laws from reaction stoichiometry!

Basic chemistry (rate laws)

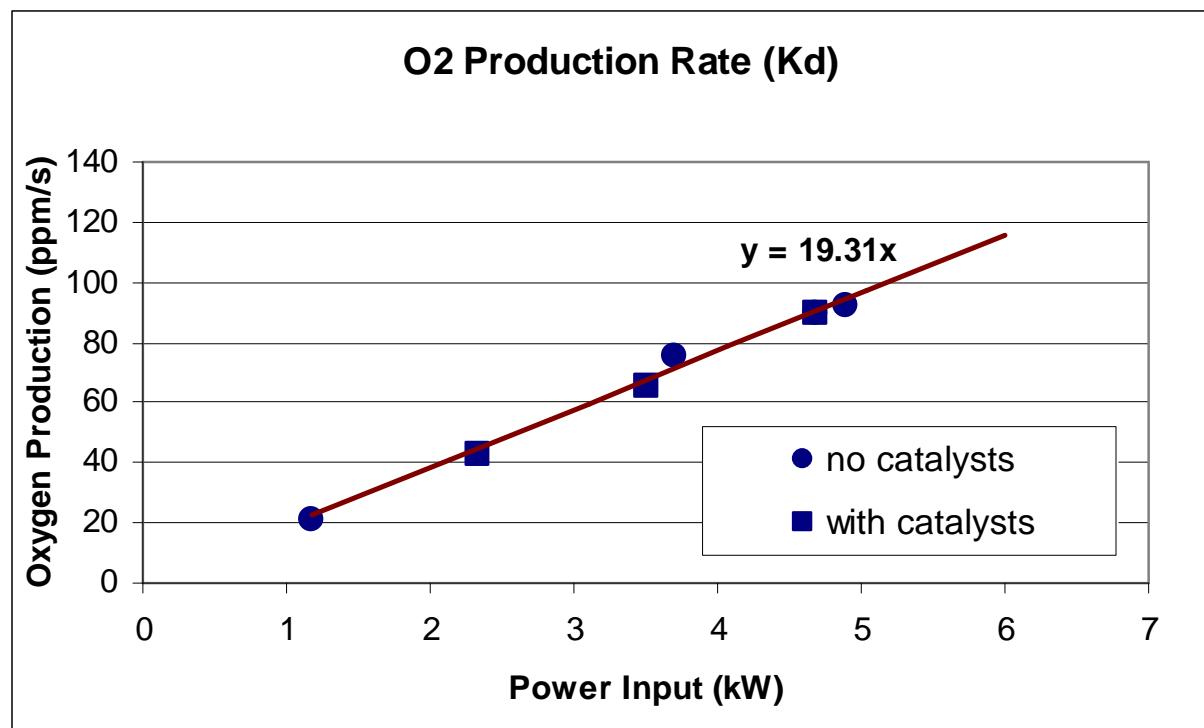
$$\frac{\partial [O_2]}{\partial t} = -k_c M_g [O_2]^n [CO]^m - k_T [O_2]^p [CO]^q + k_d$$

oxygen recombination (catalysts)

oxygen generation (discharge)

- Can we explain?
- Can we predict?
- Can we improve?

Discharge parameters



~ 90ppm/s @ 400Hz and 22.5kV

A word on mathematics

$$\dot{x}(t) = \cancel{k}_c - k_T x(t) - k_c M_g x^n(t)$$

n = 0 or 1 Already solved (easy)

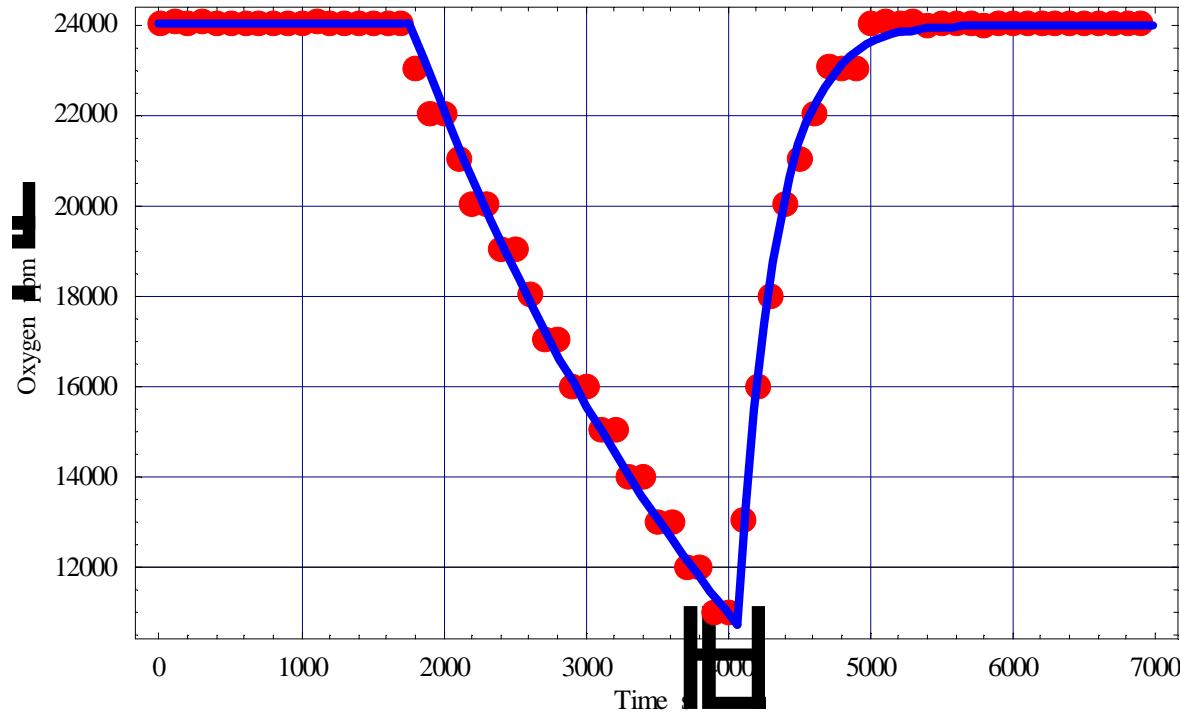
n = 2 Riccati equation (require a particular solution)

Any n? Numerically only

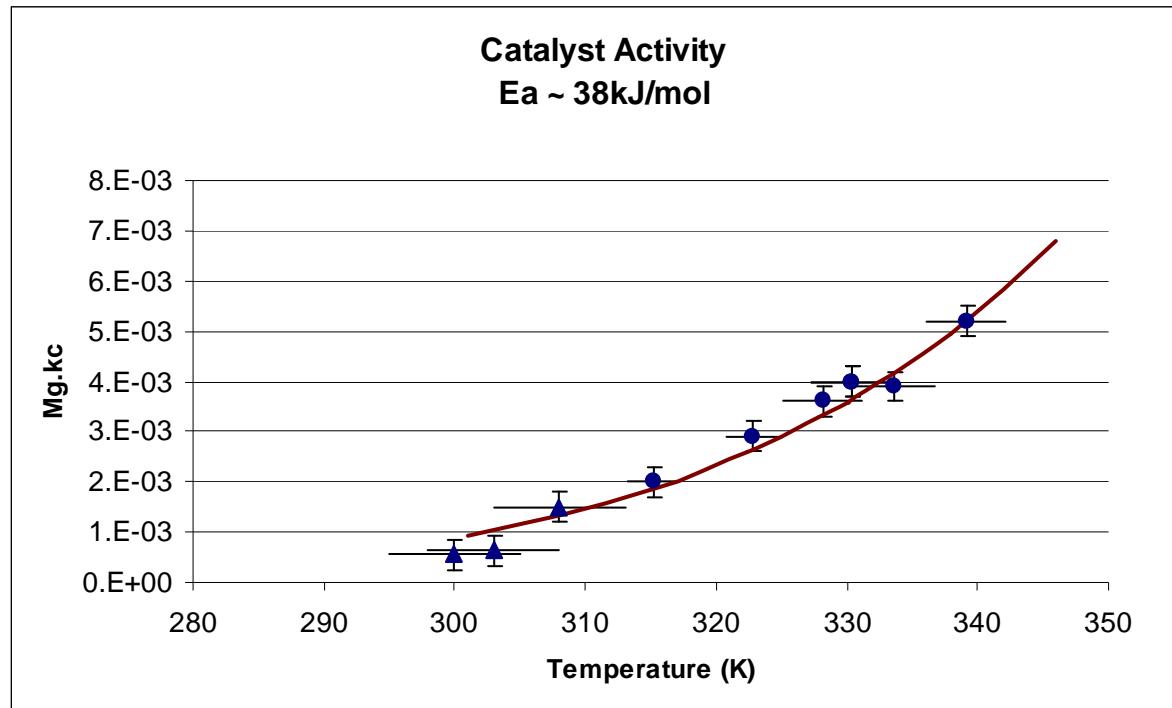
BUT, remove constant and you have an easy to solve Bernoulli equation!

With catalysts

$$\dot{x}(t) = k_d - (k_T + k_c M_g) x(t) \longrightarrow x(t) = \left(x_o - \frac{k_d}{k_T + k_c M_g} \right) \exp(-(k_T + k_c M_g)(t - t_o)) + \frac{k_d}{k_T + k_c M_g}$$



Recombination rate



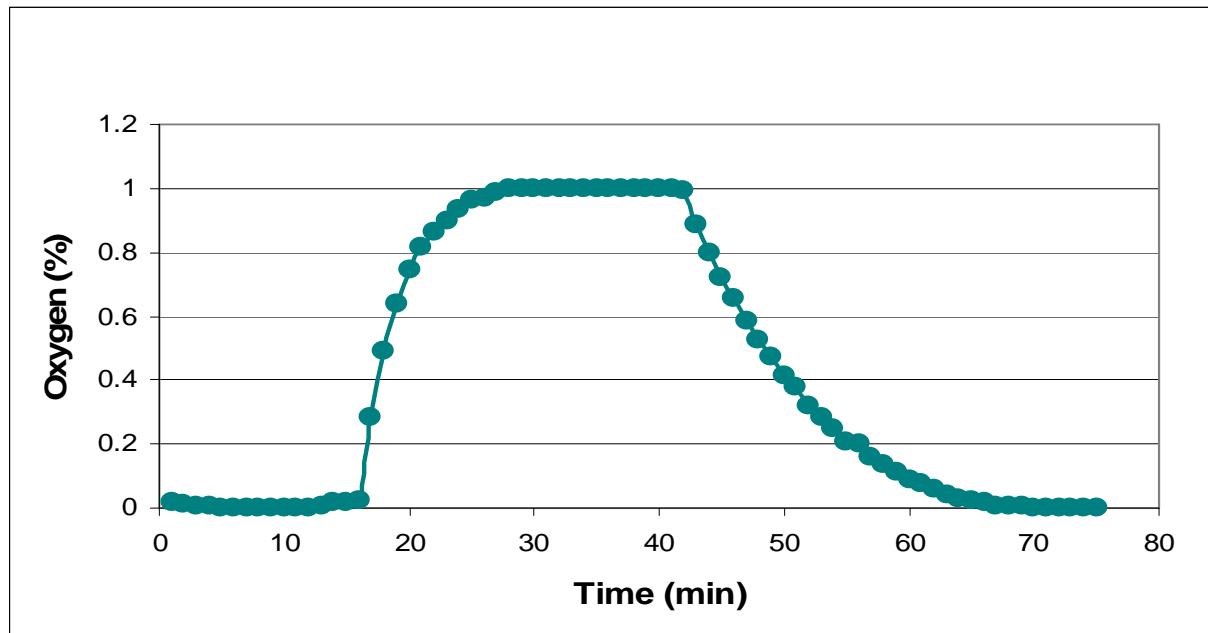
$$k_c(T) = A \exp\left(-\frac{E_a}{RT}\right)$$

$$T(T_0, P) = T_0 + \frac{RT_0}{P_0 l d v C_v} P$$

Predictions

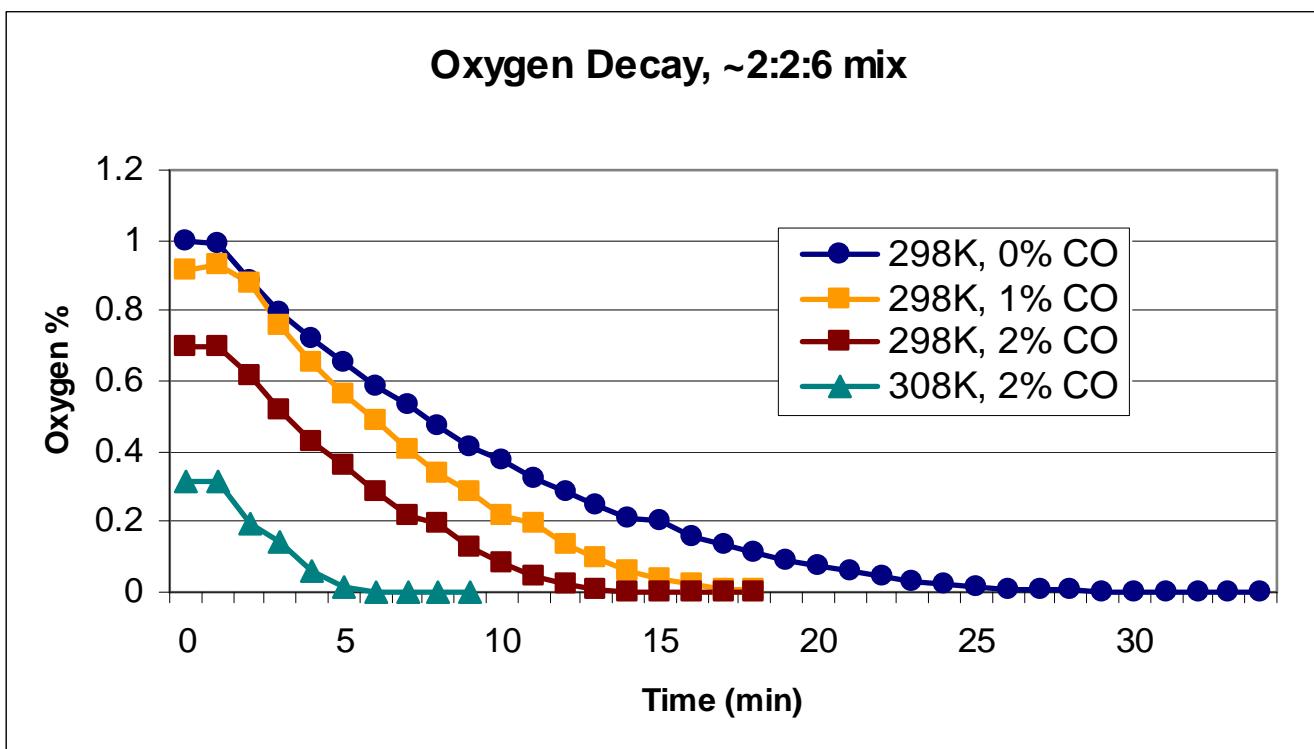
1. Oxygen levels will stabilise

$$x(t = \infty) = x_{eq} = \frac{k_d}{k_c M_g}$$



Predictions

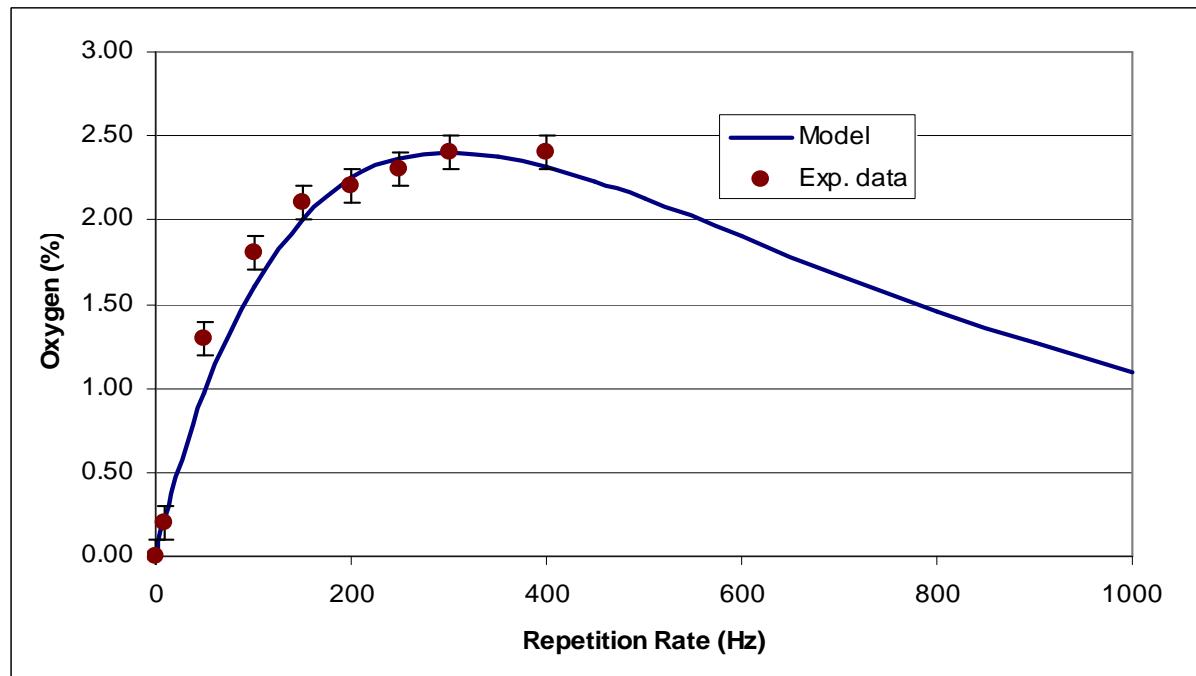
2. Increasing the initial reactant concentrations decreases the time to equilibrium; increasing the temperature will decrease equilibrium levels.



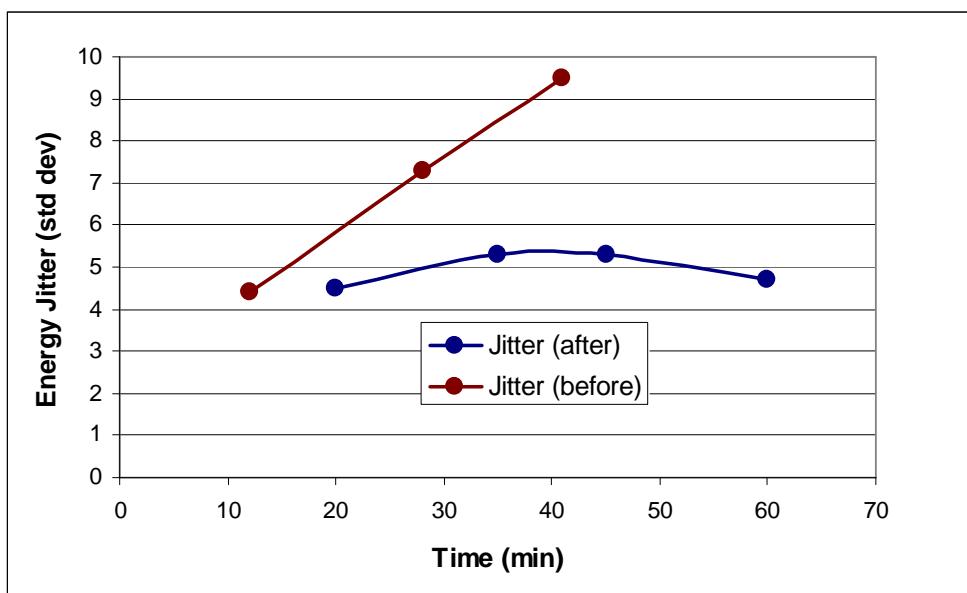
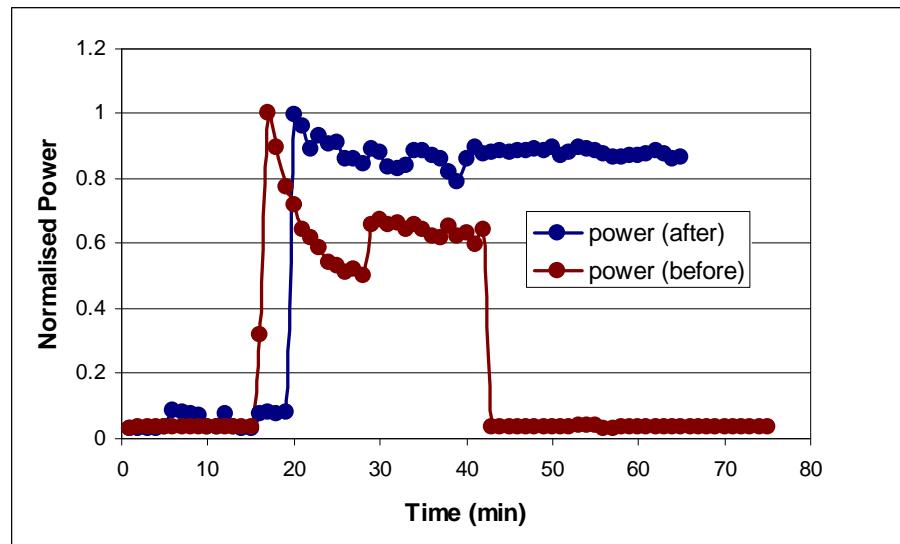
Predictions

3. The equilibrium values can decrease with increasing power!

$$P^* = \left(\frac{P_0 l d v C_v}{R T_0} \right) \cdot \frac{E_a - 2 R T_0 - \sqrt{E_a (E_a - 4 R T_0)}}{2 R}$$



Outcome

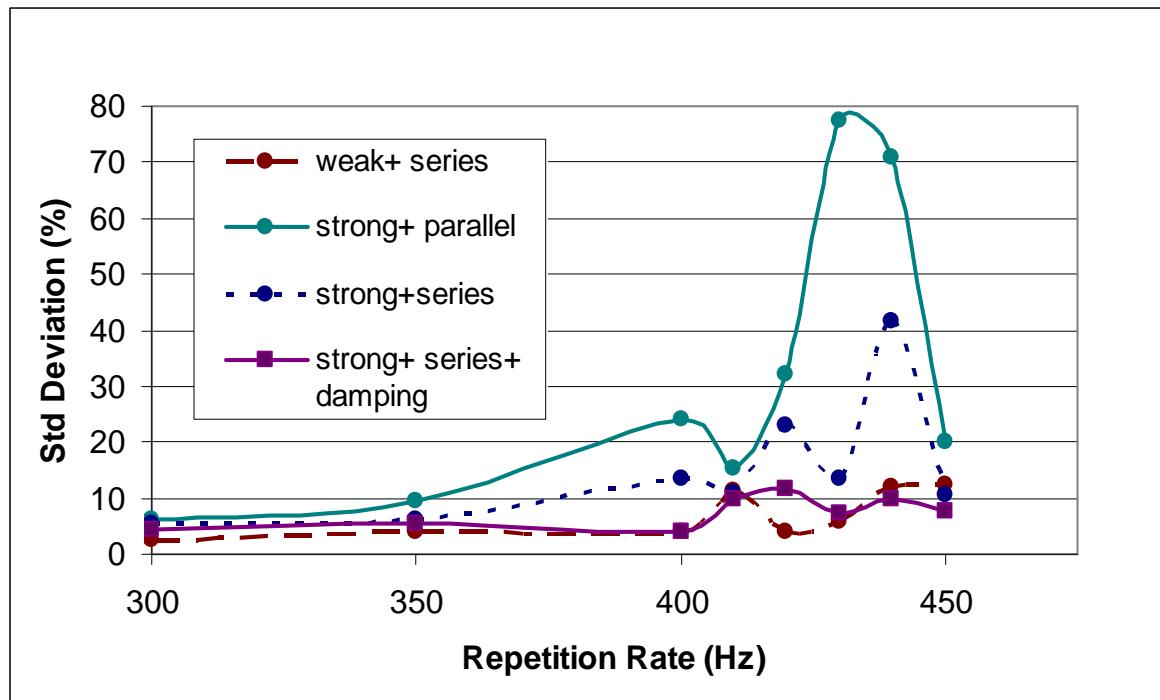


Discharge disturbances

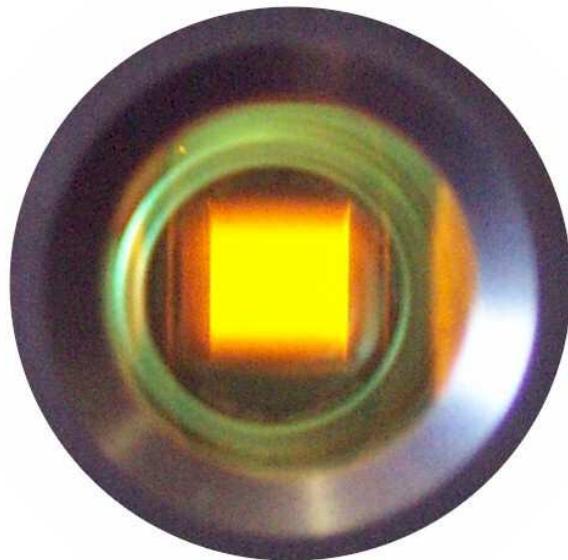
Acoustics and shocks



Acoustics



Stable discharge



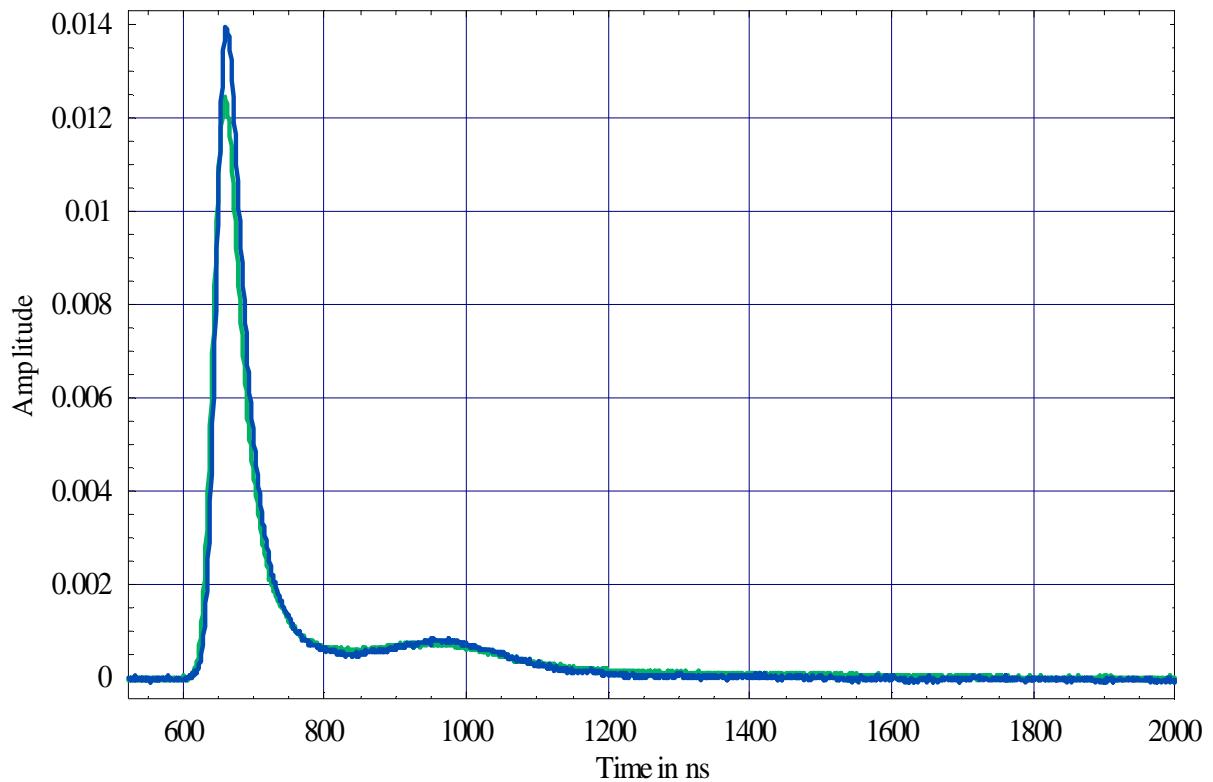
Final system

In operation at Lockheed Martin for the JSF



Final laser

- 6x improvement in efficiency
- 2x improvement in energy
- 2x improvement in production throughput



With some very interesting science along the way ...

Thank you and Acknowledgements

Co-authors:

Dr Lourens Botha (National Laser Centre)

Mr Neil du Preez (SDI)

Dr Tommy Drake (Lockheed Martin)

